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Basic Atomic Structure Worksheet H and
the 1. 2. 4. 5. 6. 8. 9. The 3 particles of
the atom are: protons (p) (e) Their
respective charges are: The number of
protons in ...

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The atomic number tells you the number of _____ in one atom of an element. It also tells you the number of _____ in a neutral atom of that element. The atomic number gives the “identity” of an element as well as its location on the periodic table.

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Group 1 elements in the periodic table are known as the alkali metals. They include lithium, sodium and potassium, which all react vigorously with air and

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water. Group 2 metals are known as the alkaline earth metals and include beryllium, barium, magnesium, calcium and strontium. ... 1-4 Vocabulary Worksheet SE - Atomic Structure (FREE) 1-4 ...

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Atomic Structure Worksheet. Label the parts of an atom on the diagram below.

4. What type of charge does a proton have?
5. What type of charge does a neutron have?
6. What type of charge does an electron have?
7. Which two subatomic particles are located in the nucleus of an atom?
8. ...

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This complete lesson covers the topic atomic structure and isotopes in terms of protons, neutrons and electrons and features a 29 slide interactive PowerPoint including starter, group activities, questions and plenary along with an 8 page worksheet and a lesson

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plan. An exam style question is also included. Answers to all exercises are provided.

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After filling the first five rows, we still have $80 - 54 = 26$ more electrons to accommodate. According to Figure 1.3.1, we need to fill the 6s (2 electrons),

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4f (14 electrons), and 5d (10 electrons) orbitals. The result is mercury's electron configuration: $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10} = \text{Hg} = [\text{Xe} \dots$

1.4: Atomic Structure - Electron Configurations ...

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It has two valence electrons but is in Group 1. Period 2, group 1 element. It has more than 50 but less than 80 electrons and is in Group 17. This element is a transition metal that is liquid at room temperature. It has the lowest atomic number of any Group 7 elements. Period 4, alkali metal. It has a higher atomic number than carbon and lower than P.

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Table 9.

The Atom for Middle School

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an atom on the diagram below. 4. What type of charge does a proton have? 5. What type of charge does a neutron have? 6. What type of charge does an electron have? 7. Which two subatomic particles are located in the nucleus of an atom? 8.

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